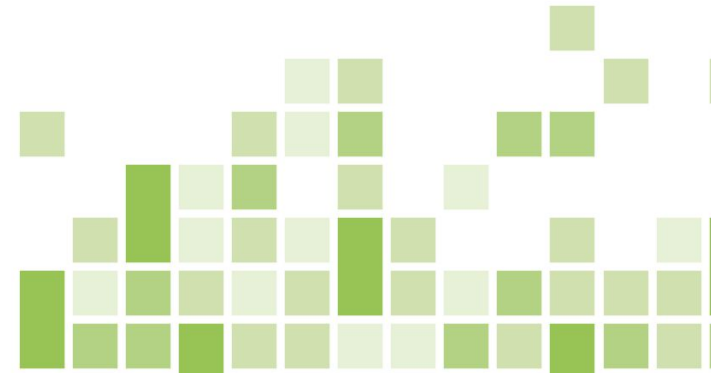




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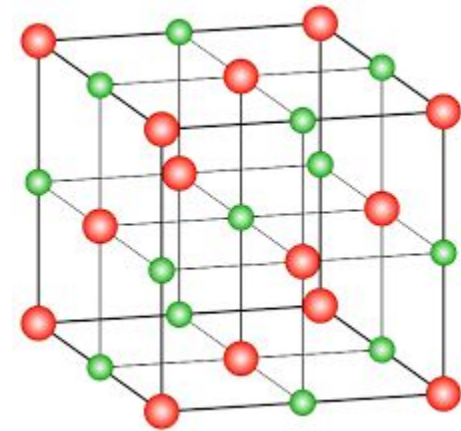
DEFECT IN SOLID

2017

What is a defect in a solid?

Solids consist of crystal. Crystals are described by scientists using an idealized model, but such crystals are not found in nature. The deviation of the crystal structure from the periodic crystal structure is defects. Crystals described by scientists using an idealized model, but such crystals are not found in nature.

The real crystal are not perfect. A real crystal always has a large number of imperfections in the lattice. One can reduce crystal defect considerably, but can never eliminate them entirely.



Model of perfect crystal



Defects are divided into four classes,
if we consider the geometric of
defects

point defects

Line defects

Surface defects

Volume defects

Defects occur in a solid at any
temperature. The number of defects
increases with increasing
temperature and under the influence
of ionizing radiation and
mechanical treatment



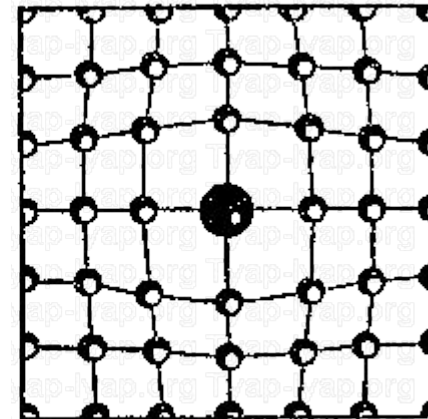
Point defects include vacancies, atoms between the nodes of the crystal lattice, impurities.

Vacancy – An atom missing from regular lattice position. Vacancies are present invariably in all materials.

An **impurity** is a foreign atom that does not form a chemical compound in the crystal lattice



Point defect-vacancy

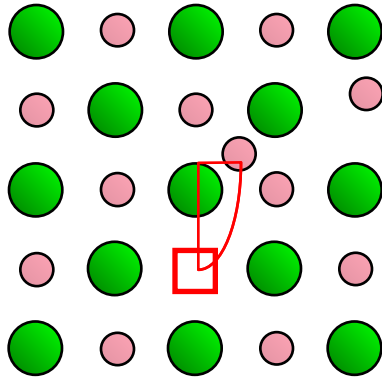


Point defect-impurity

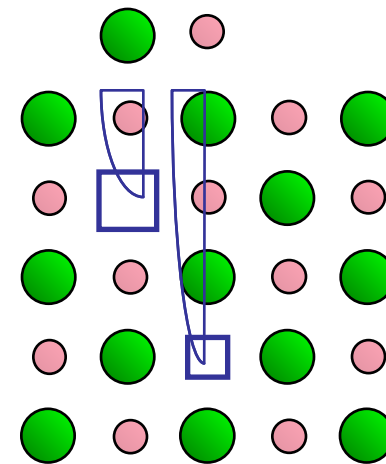
When the temperature is sufficiently high, as the atoms vibrate around their regular positions, some acquire enough energy to leave the site completely. When the regular atom leaves, a vacancy is created.

A pair of one cation and one anion can be missed from an ionic crystal. Such a pair of vacant ion sites is called *Schotky defect*.

In crystals, an ion displaced from a regular site to an interstitial site is called *Frenkel defect*.



**Frenkel defect: anion
vacancy-interstitial cation pair**



**Schottky defect: anion
-cation vacancy pair**

- The number of vacancies (defect of Schotky) is given by formula:

$$n = N \exp \left[- \frac{E_v}{k_B * T} \right]$$

- N-number of atom;
- n- number of vacancies;
- k_B - constant of Boltzman, $1,38 * 10^{-23}$ J /K;
- T- temperature of solid;
- E_v - energy of formation defect of Schotky;

- The number of defect of pair Frenkel is given by formula:

$$n = (N' N)^{1/2} \exp\left[-\frac{E_f}{2k_B * T}\right]$$

- N-number of atom;
- N' - number of interstitial by can take up free atom;
- n- number of defects pair of Frenkel;
- k_B - constant of Boltzman, $1,38 \cdot 10^{-23}$ J /K;
- T- temperature of solid;
- E_f - energy of formation pair of Frenkel.

- The experimental data do not show that the observed values of the yield strength are much smaller than the theoretical values. To explain this effect, a new two defect-dislocation.

The two types of dislocations are:

- Edge dislocation
- Screw dislocation
- The experimental data do not show that the observed values of the yield strength are much smaller than the theoretical values. To explain this effect, a new type of defect- dislocation.

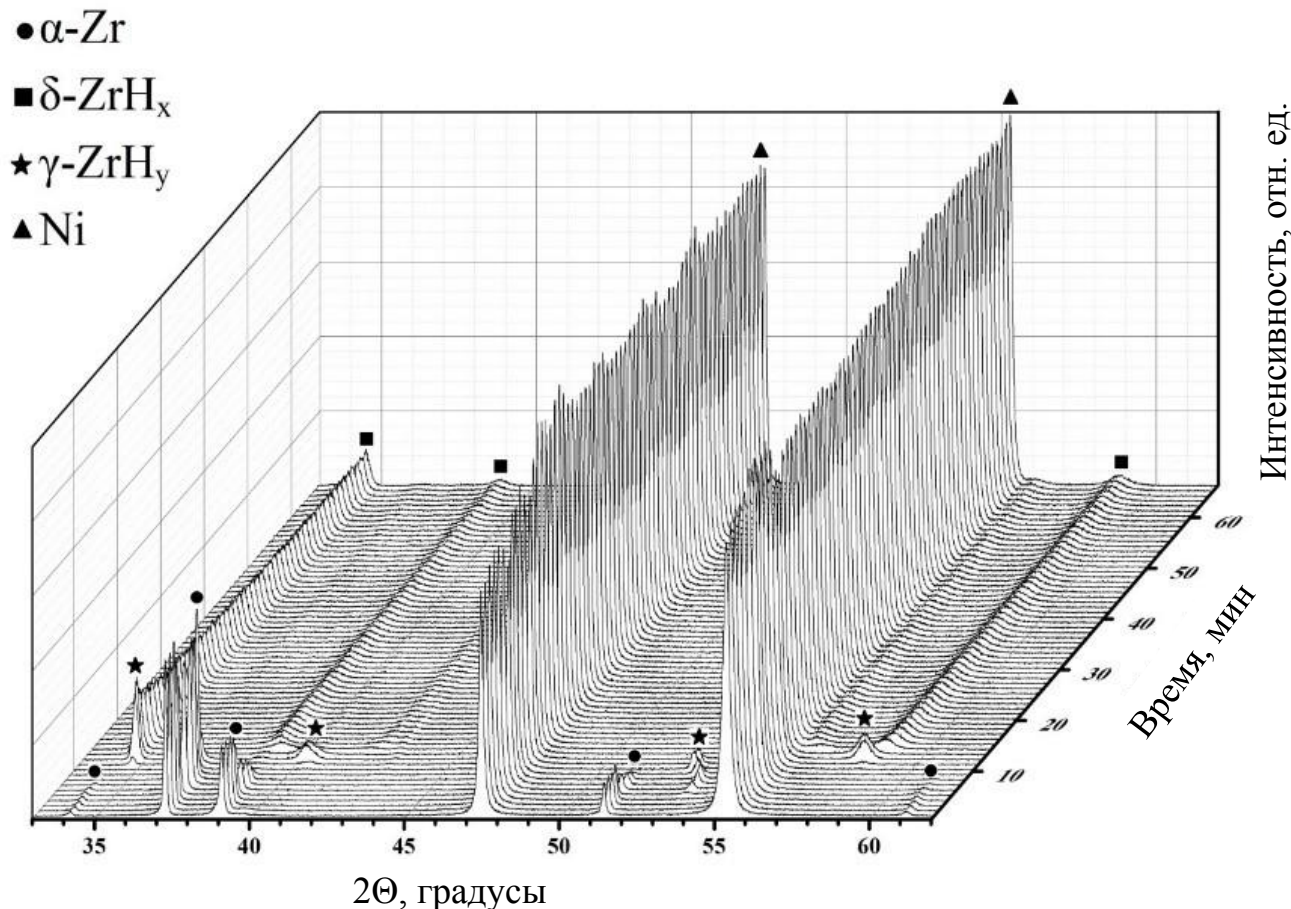


Рисунок 4 – Фазовые переходы в системе цирконий-водород с никелевым покрытием в процессе насыщения водородом при температуре 550 °С

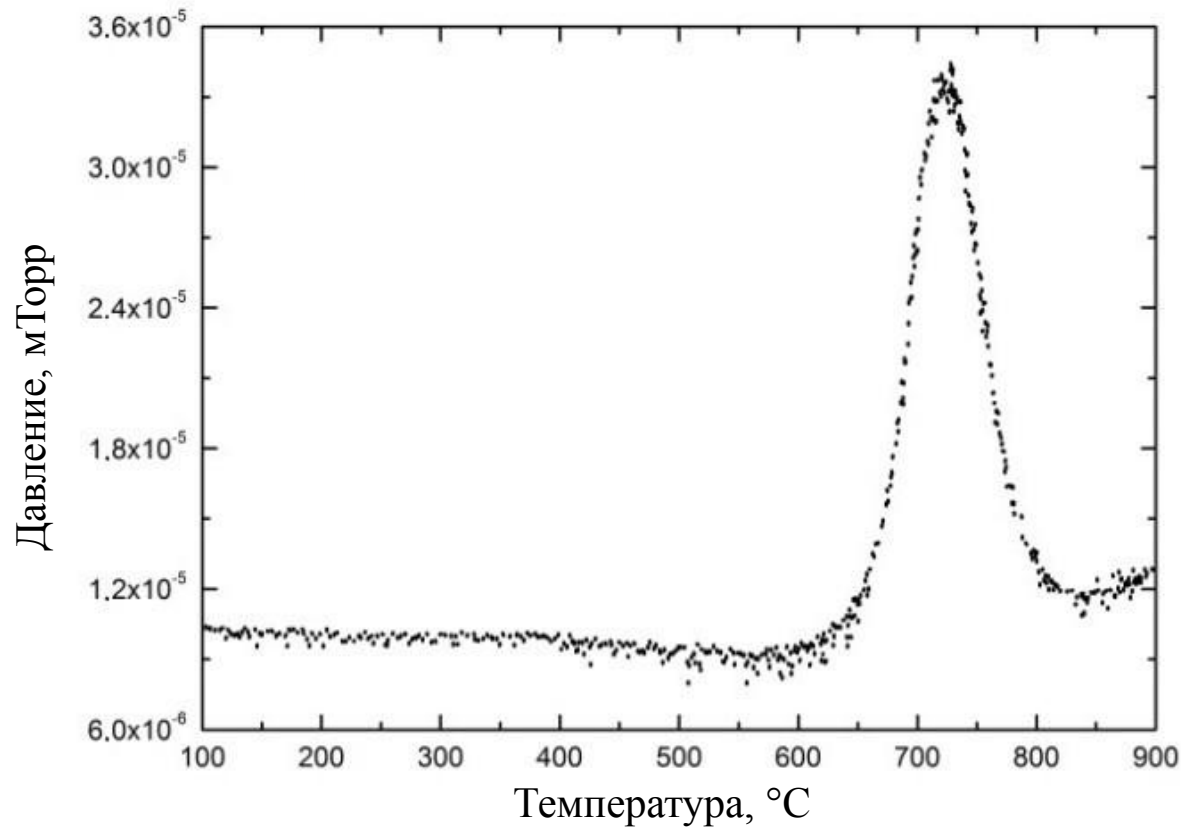


Рисунок 5 – Спектр термостимулированной десорбции водорода из образца циркониевого сплава после гидрирования в газовой атмосфере



Исследовано взаимодействие водорода с циркониевым сплавом Zr1%Nb с нанесенным слоем никеля при газофазном гидрировании при температурах 350, 450 и 550 °С. Были сделаны следующие выводы:

- Скорость поглощения водорода циркониевым сплавом Э110 увеличивается в 16,8 раза при повышении температуры;
- При температуре 350 °С гидрирование приводит только к фазовому переходу α -циркония к δ -гидриду циркония, тогда как при 450 и 550 °С процесс протекает с образованием метастабильной тетрагональной фазы гидроксида циркония;
- При температуре 522-540 °С наблюдается процесс диссоциации гидридов в циркониевом сплаве Э110;
- При дальнейшем повышении температуры происходит увеличение β -фазы циркония из-за переходной фазы $\alpha \rightarrow \beta$ циркония.



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