Шаблон для презентации

Now the end carbon was a single unprised electry. What we do with it? Before the bond broke, the end carbon was called all (sp² hybridized). We might think that the single electron would still be in an sp² orbital. However, since an sp² orbital cannot overlap efficiently with a π bond, the single electron would be have to be localized on the end carbon atom. If the end carbon atom becomes trigonal (sp² hybrid to the single electron could be in a p orbital and this could overlap and combine with the π by within would mean that the radical could be spread over the molecule in the same orbits what core in the cation.

Solonce against have the orbitals to

has me more tooms, as same words, and so the same energy levels. In fact, the molecular orbital dergy level agram orbits compound is almost the same as the one for the allyl cation: the only difference as the number of electrons in the π system. Whereas in the allyl cation π system we only had two electrons, here we have three (two from the π bend plus the single one). Where does this extra electron go? Answer: in the next lowest molecular orbital—the nonbonding molecular orbital.

efficient overlap of

Вы можете использовать данное оформление для создания своих презентаций. Финк Вероника Георгиевна, Костанайская область

nis MO now has one electron in it. It's known as the Singly Occupied Albiecular Orbital (SOLO)







