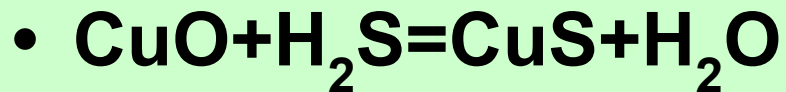
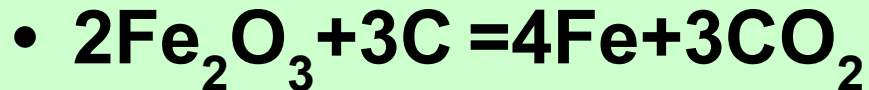
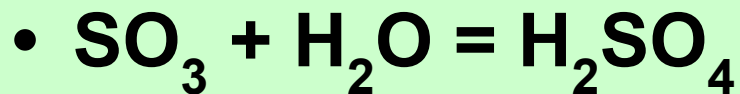
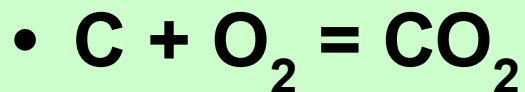


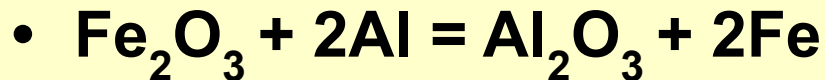
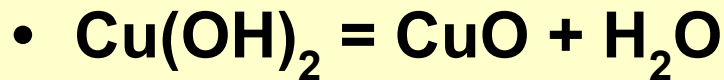
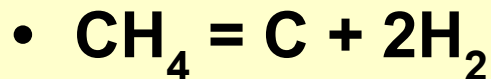
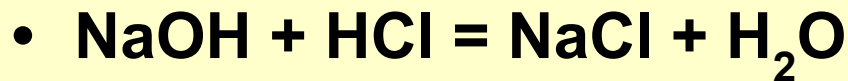
Составление уравнений ОВР

Какие из реакций являются окислительно-восстановительными? Укажите окислитель и восстановитель

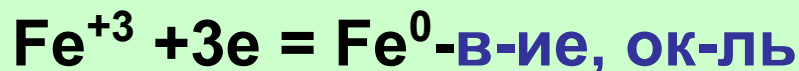
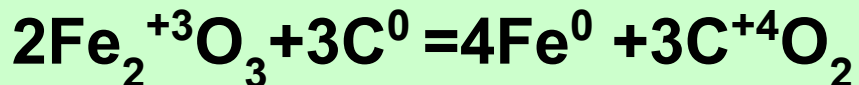
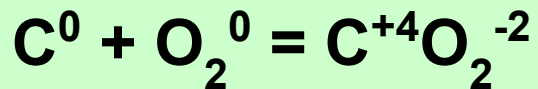
• 1 Вариант



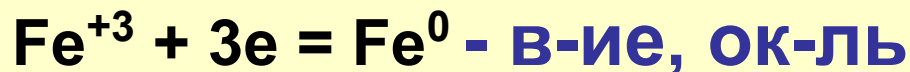
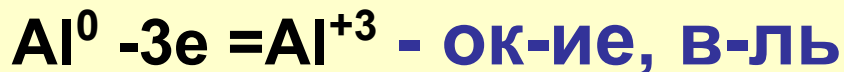
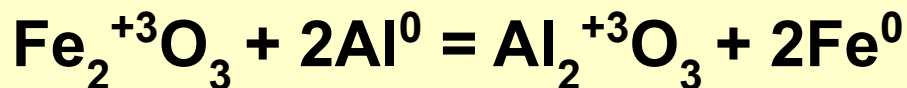
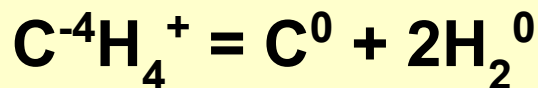
• 2 Вариант

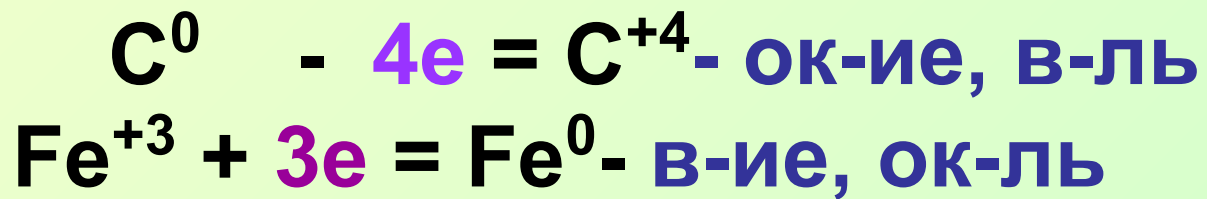
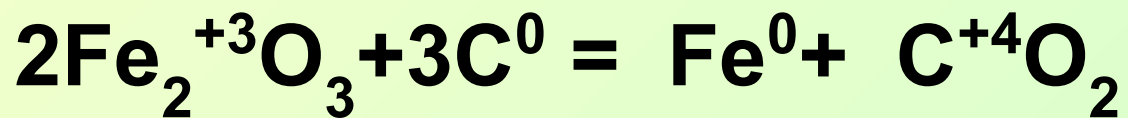


- 1 Вариант



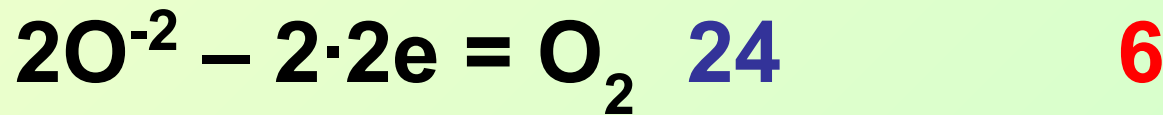
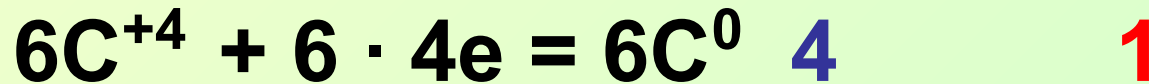
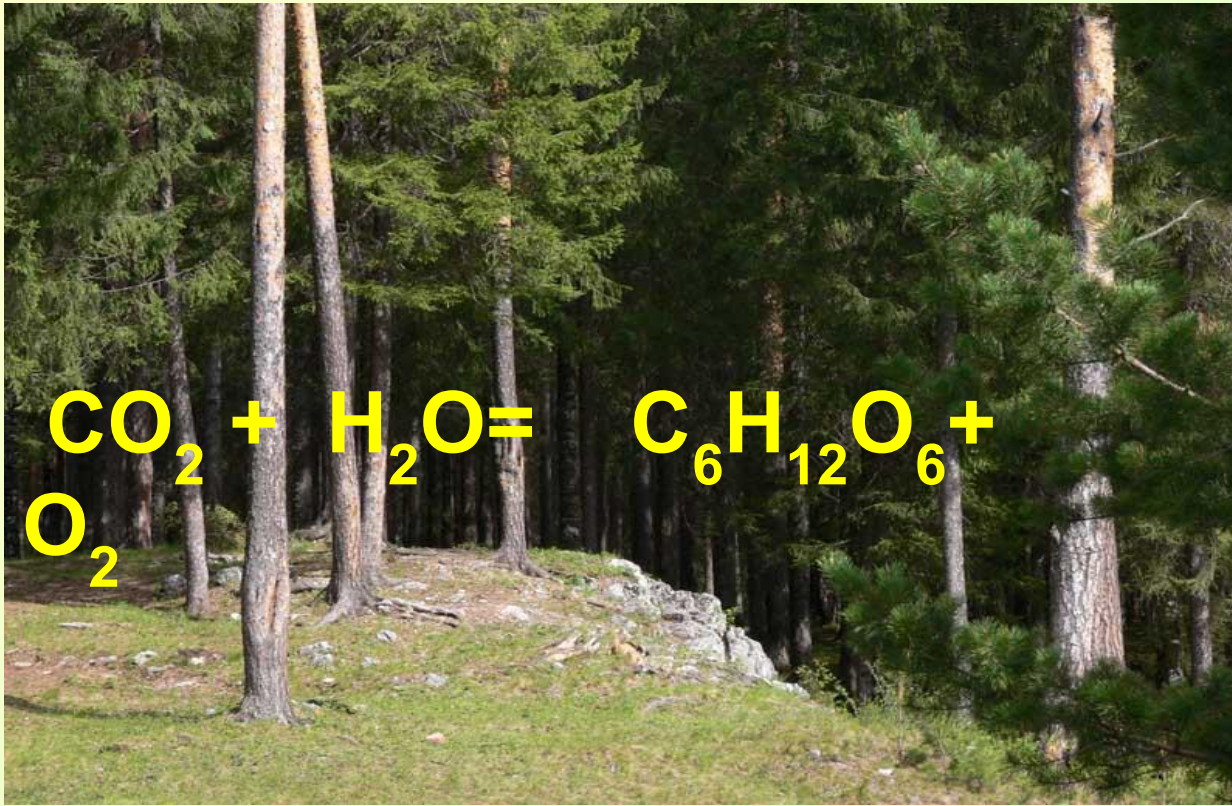
- 2 Вариант

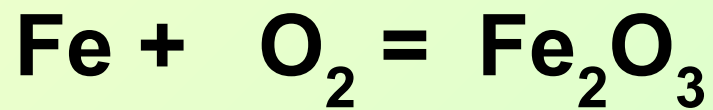




3

4





$\text{Fe}^0 - 3e = \text{Fe}^{+3}$ - ок-ие, в-ль

$\text{O}_2 + 4e = 2\text{O}^{-2}$ - в-ие, ок-

ль

4

3

Расставьте коэффициенты методом электронного баланса

- $\text{Zn} + \text{HCl} = \text{ZnCl}_2 + \text{H}_2$
- $\text{Cu} + \text{HNO}_3 = \text{Cu}(\text{NO}_3)_2 + \text{NO}_2 + \text{H}_2\text{O}$
 - $\text{Al} + \text{O}_2 = \text{Al}_2\text{O}_3$
 - $\text{P} + \text{O}_2 = \text{P}_2\text{O}_5$

Домашнее задание

- § 43
- Раб. тетр.: Стр. 162-163, упр. 12
упр. 13(а-в)
упр. 13